

## Weighing In HW

Read and outline (Cornell style) **Sections 11.2 AND 11.3** in your chemistry textbook. Then answer the following assessment questions. **\*\*Your section outline must be at least ½ page and have 3-4 questions.\*\***

1. Explain what is meant by molar mass.
2. Set up a bridge (or bridges) to convert the mass of an element to the number of atoms (particles) of the element. (Hint: Just set up the bridges with the appropriate units on top and bottom. Don't worry about any numbers.)
3. Describe how you can determine the molar mass of a compound, like  $\text{CH}_3\text{OH}$ .
4. Calculate the number of moles in each of the following masses:
  - a. 3.0 g of  $\text{BBr}_3$
  - b. 111.8 g of Fe
  - c. 0.00655 g of Sb
  - d. 89.0 g of NaCl
5. Calculate the mass of each of the following amounts:
  - a. 7.0 mol of Ti
  - b. 1.002 mol of Cr
  - c. 0.005 mol of  $\text{C}_3\text{H}_8$
  - d. 0.120 mol of  $\text{Ca}(\text{OH})_2$
6. Calculate the number of particles (atoms or molecules) in each of the following:
  - a. 0.431 g of LiF
  - b. 12.5 g of  $\text{C}_6\text{H}_{12}\text{O}_6$
  - c. 550 mol of Al
  - d. 7.8 mol of  $\text{BaCl}_2$